

# Displacement of Metals

## Chemical 16 Equations

### General Rule



A metal will be displaced from a salt by a **more reactive** metal.

e.g. 1. Magnesium + Lead Oxide  $\rightarrow$  Magnesium Oxide + Lead

A metal will not be displaced from a salt by a **less reactive** metal.

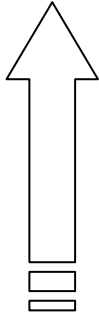
e.g. 2. Lead + Magnesium Oxide  $\rightarrow$  No Reaction

### Task 1

**In your exercise book**, write word equations for the following reactions. If you have been taught how to, then write a balanced symbol equation under each word equation (assume a valency of 3 for iron, 2 for lead and 1 for copper).

	Reactants
a.	Aluminium and lead chloride
b.	Potassium and calcium chloride
c.	Sodium and iron chloride
d.	Zinc and aluminium chloride
e.	Lead and copper chloride
f.	Iron and copper oxide
g.	Calcium and lead iodide
h.	Tin and zinc sulphate
i.	Calcium and copper sulphate
j.	Sodium and potassium fluoride
k.	Magnesium and lead nitrate
l.	Iron and zinc sulphate
m.	Copper and aluminium carbonate
n.	Potassium and magnesium chloride
o.	Zinc and lead sulphate
p.	Magnesium and calcium iodide
q.	Aluminium and lead bromide
r.	Potassium and sodium carbonate
s.	Calcium and copper nitrate
t.	Zinc and copper sulphate

### Reactivity Series

<b>most reactive</b>  <b>least reactive</b>	Potassium
	Sodium
	Calcium
	Magnesium
	Aluminium
	Zinc
	Iron
	Tin
	Lead
	Copper